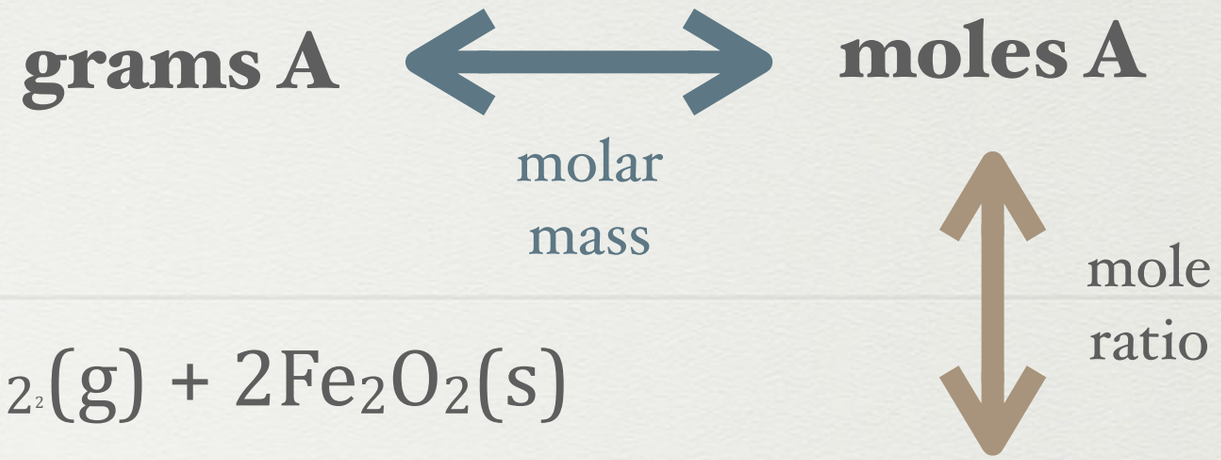
Warm Up 3/5/2015

$$4FeS_2(s) + 6\,O_2 \rightarrow 4\,SO_{2_2}(g) + 2Fe_2O_2(s)$$

* For 4.2 moles of SO_2, how many grams of O_2 will there be?



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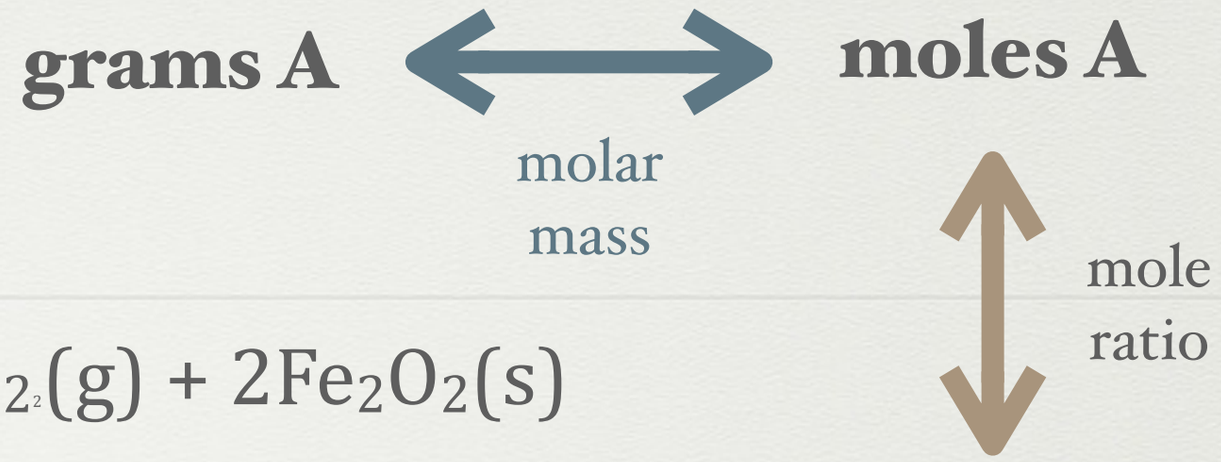
Assess the problem:

What value is given?

What value are you solving for?



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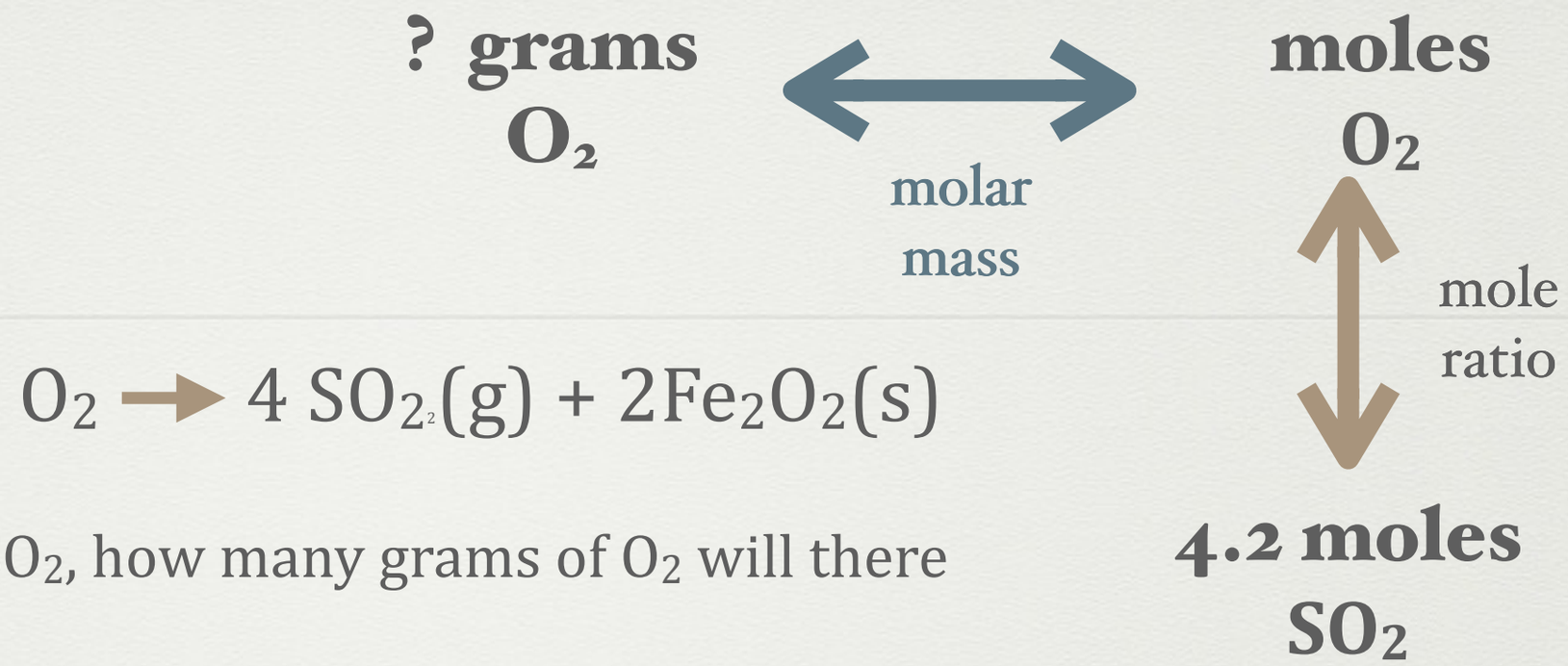
Given: 4.2 moles of SO_2

Solving for: grams of O_2

Identify start and end points on map:

- **Starting point is what's given.**
- **Ending point is what you are solving for.**



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Setup your dimensional analysis

$$\frac{4.2\ \text{moles } SO_2}{} \left| \frac{6\ \text{moles } O_2}{4\ \text{moles } SO_2} \right| \frac{32\ \text{grams } O_2}{1\ \text{mole } O_2} = ?\ \text{grams } O_2$$