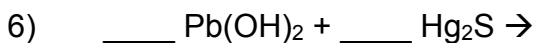
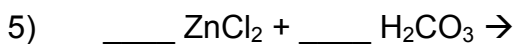
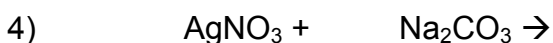
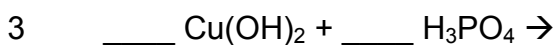
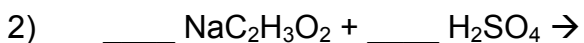
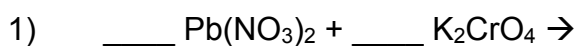


Reaction Products Worksheet

Name _____

For each of the following reactions, determine what the products of each reaction will be. When you have predicted the products, balance the equation and use a table of solubility products to determine which of the products (if any) will precipitate. Assume all reactions take place in water.

Write the balanced overall equation and the net ionic equation if a precipitate is formed.



SOLUBILITY RULES WORKSHEET

Classify each of the following as soluble (aq) or insoluble (s).

1) FeCl_3

2) KOH

3) KCl

4) $\text{Fe}(\text{OH})_3$

5) $(\text{NH}_4)_2\text{SO}_4$

6) AgNO_3

7) PbSO_4

8) $\text{Mn}(\text{OH})_2$

9) NaBr

10) AgCl

11) NaNO_3

12) BaSO_4

Classify each of the substances as being soluble or insoluble in water.

KBr =

PbCO_3 =

BSO_3 =

zinc hydroxide =

sodium acetate =

silver iodide =

cadmium (II) sulfide =

zinc carbonate =

silver acetate =

copper (II) sulfide =

$\text{Mg}_3(\text{PO}_4)_2$ =

KOH =

NiCl_2 =

NH_4OH =

Hg_2SO_4 =

PbI_2 =

Reaction Products Worksheet - Key

For each of the following reactions, determine what the products of each reaction will be. When you have predicted the products, balance the equation and use a table of solubility products to determine which of the products (if any) will precipitate. Assume all reactions take place in water.



copper (II) phosphate precipitates

