

Reduction - Oxidation "Redox Reactions"

Reduced: Gains electrons

Oxidizing agent: accepts
the electrons

Oxidized | Loss of electrons

Reducing agent: Donates
the electrons

Oxidized

Is

Loss

Reduction

Is

Gained

Oxidation numbers

1.) An ion charge is the oxidation number



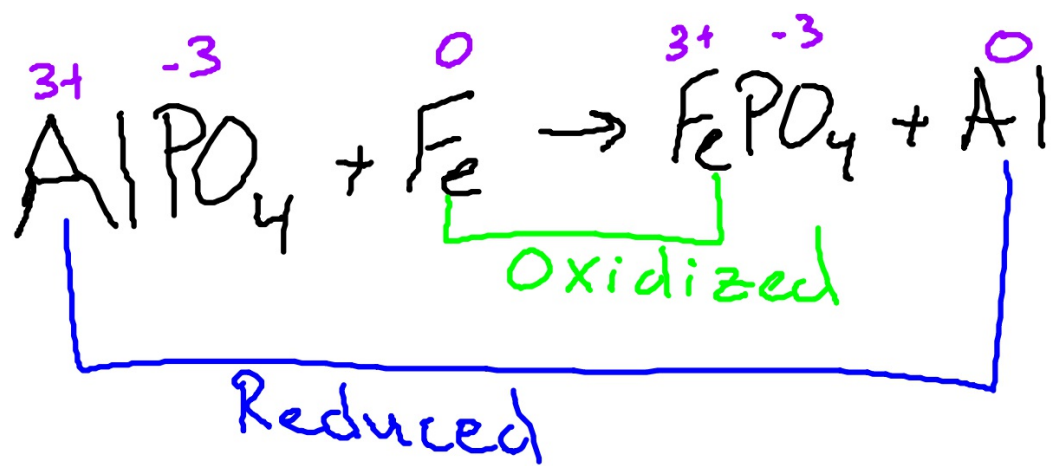
2. All compounds must
equal zero

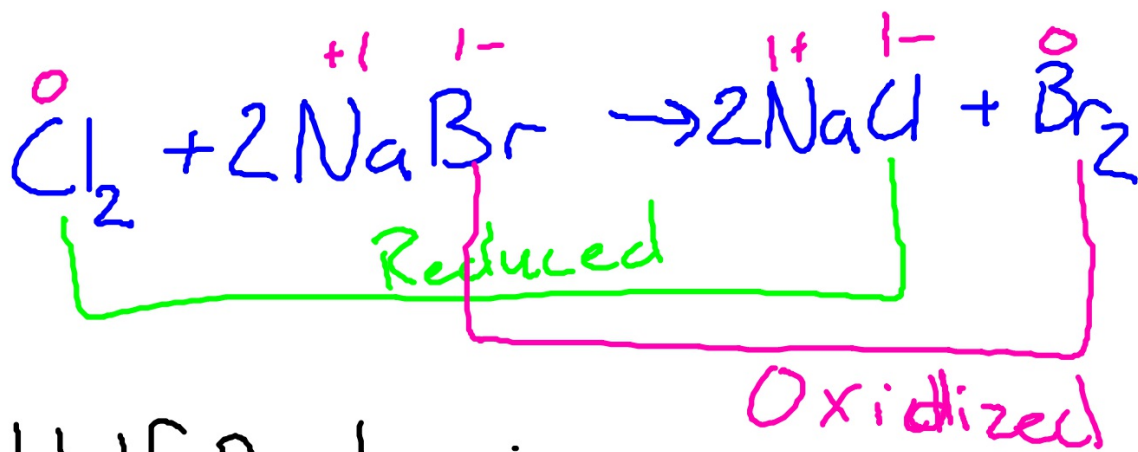
- Oxygen is -2 (usually)

- Hydrogen is $+1$ (mostly)

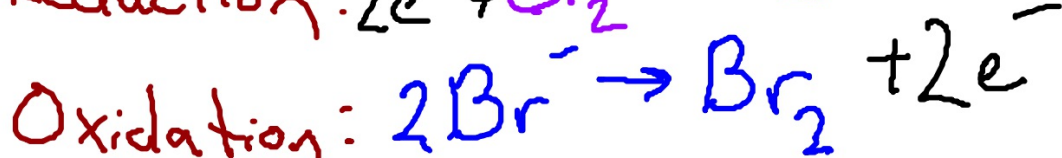
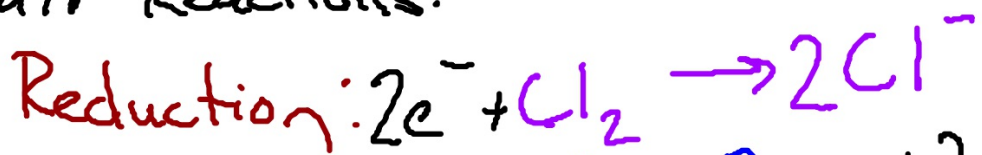
3.) All elements are zero

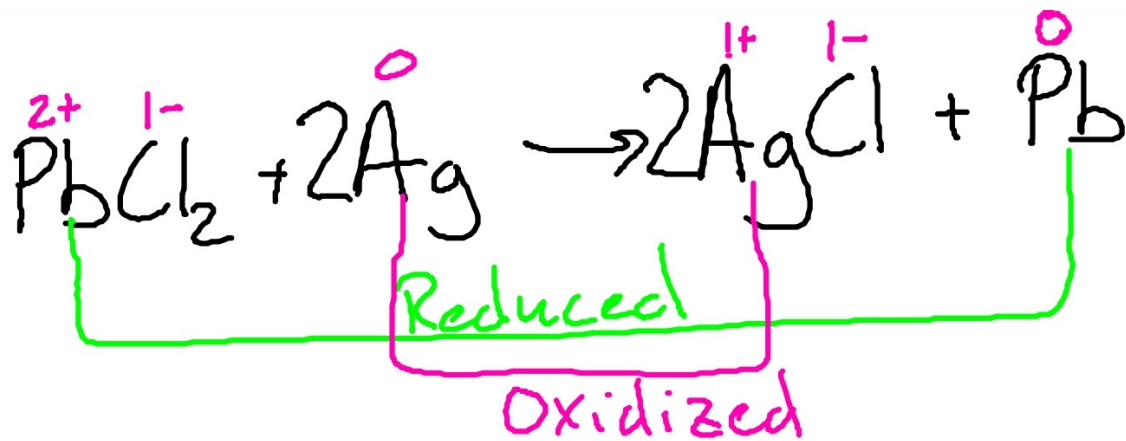
Ex: Fe, Na, H₂, N₂



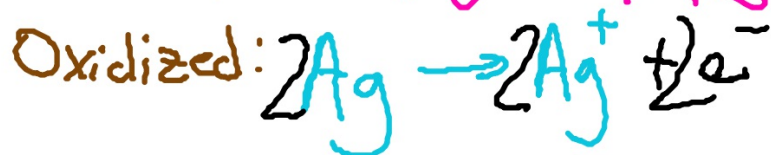
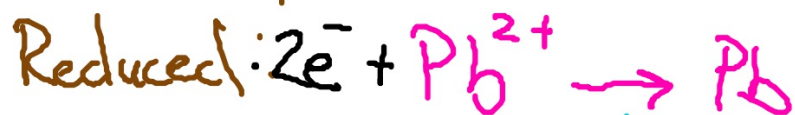


Half Reactions:

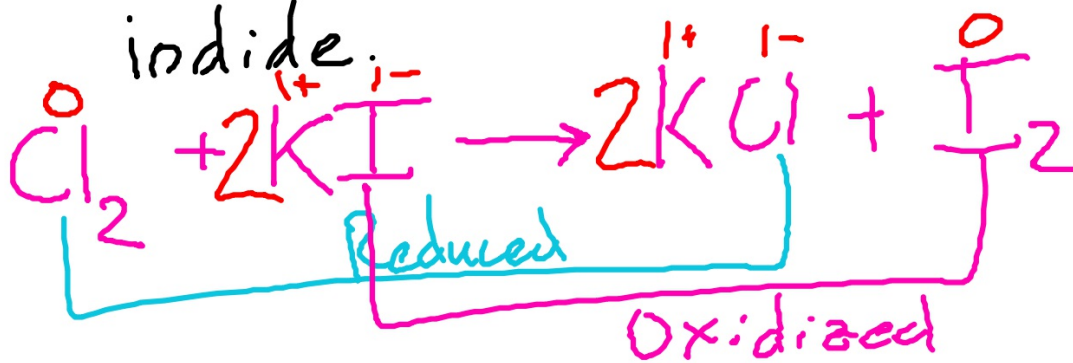




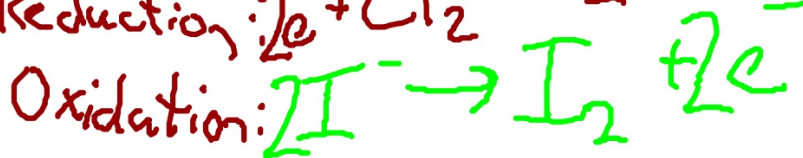
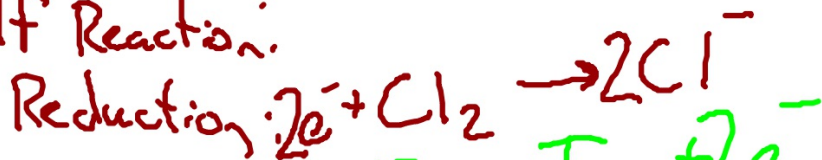
Half Rxn



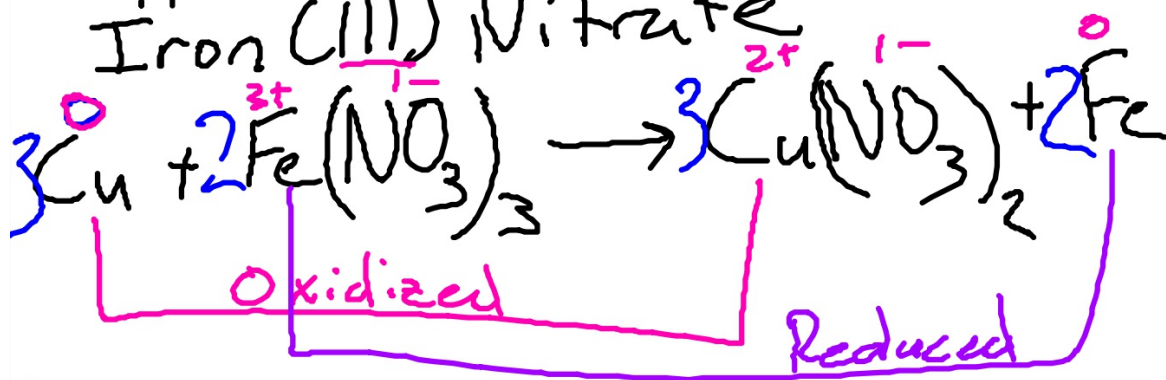
Chlorine reacts with potassium
iodide.



Half Reaction:



Copper metal reacts with
Iron (III) Nitrate



Half Rxn:

