

Conversion from
g \rightarrow mol

12.5g Fe

Fe - 55.85g/mol

$$12.5\cancel{\text{g}} \text{ Fe} \times \frac{1 \text{ mol}}{55.85\cancel{\text{g}}} = 0.223 \text{ mol Fe}$$

348g $\text{Mn}_3(\text{PO}_4)_4$

$$348\cancel{\text{g}} \text{Mn}_3(\text{PO}_4)_4 \times \frac{1 \text{ mol}}{544.7\cancel{\text{g}}}$$

0.638 mol $\text{Mn}_3(\text{PO}_4)_4$

82.14 grams of Beryllium Fluoride



$$\text{Be} - 9.01 \times 1 = 9.01$$

$$\text{F} - 19.00 \times 2 = 38.00$$

$$\underline{47.01 \text{ g/mol}}$$

$$82.14 \text{ g BeF}_2 \times \frac{1 \text{ mol}}{47.01 \text{ g}} =$$

$$1.74 \text{ mol BeF}_2$$

mol \rightarrow g

2.5 mol H_2SO_4

$$\begin{array}{r} H - 1.0079 \times 2 = 2.014 \\ S - 32.07 \times 1 = 32.07 \\ O - 16.00 \times 4 = 64.00 \\ \hline 98.08 \end{array}$$

$$2.5 \text{ mol } H_2SO_4 \times \frac{98.08 \text{ g}}{1 \text{ mol}} =$$

$$245.2 \text{ g } H_2SO_4$$

$$98.08 \text{ g/mol}$$

$$7.25 \text{ mol GeS}_2 \quad \begin{array}{l} \text{Ge} - 72.61 \times 1 = 72.61 \\ \text{S} - 32.07 \times 2 = 64.14 \\ \hline 136.75 \end{array}$$

$$7.25 \text{ mol GeS}_2 \times \frac{136.75 \text{ g}}{1 \text{ mol}} =$$

$$991.43 \text{ g GeS}_2$$