Name _____

Honors Chemistry Unit 1 Review Questions
Test: Sept 21st and 22nd

SECTION 3.1 MEASUREMENTS AND THEIR UNCERTAINTY

Using different rulers, Bruce and Pete each measure the length of the same object three times.

1. Bruce's three measurements are 19 cm, 20 cm, and 22 cm. Calculate the average value of his measurements and express the answer with the correct number of significant figures.
2. Pete's three measurements are 20.9 cm, 21.0 cm, and 21.0 cm. Calculate the average value of his measurements and express the answer with the correct number of significant figures.
3. Multiply the answer to problem 1 by the answer to problem 2. Express the answer in scientific notation with the correct number of significant figures.
4. Whose measurements are more precise?
5. The actual length of the object is 20 cm. Whose measurements are more accurate?

8. Four boards each measuring 1.5 m are laid end to end. Multiply to determine the combined length of the boards, expressed with the correct number of significant figures.

SECTION 3.2 THE INTERNATIONAL SYSTEM OF UNITS (SI)

A fish tank measures 0.40 meter long by 0.20 meter wide by 0.30 meter high.

1. What is the width of the tank in centimeters?
2. What is the length of the tank in millimeters?
3. What is the volume of the tank in liters?
4. What is the mass of water, in grams, that would fill the tank halfway?

SECTION 3.3 CONVERSION PROBLEMS

1. The population of San Francisco is 750,000 in an area of 49 square miles. What is the population density in San Francisco? Express your answer in people per acre. (1 mi^2 = 640 acres)
2. A sugar-free powdered drink mix sells for $2.99 per can. Each can of the mix contains 50.2 g of powder, which, when added to water, will make 8 quarts of drink. What is the cost of the powdered drink mix in dollars/lb? (454 g = 1 lb)
3. A car is travelling at 60 miles per hour. Express this speed in kilometers per hour (km/h). (1 mi = 1.609 km)

6. The speed limit on a certain highway is 72 km/h. What is this speed in cm/s?
7. Gold has a density of 19.3 g/cm^3. What is the mass, in kilograms, of one cubic meter of gold?
8. An automobile can travel 40.0 miles on one gallon of gasoline. How many kilometers per liter is this? (1.61 km = 1 mi; 1 L = 0.264 gal)
9. Suppose that gold is selling at $375/ounce. How many milligrams of gold could you buy for one cent? (16 oz = 1 lb; 1 lb = 454 g)

SECTION 3.4 DENSITY

Use the data in Table 3.7 to solve problems 1–4.

1. What is the mass at 20°C of 5 liters of air?
2. A balloon filled with air is released in a room filled with carbon dioxide. Will the balloon float to the ceiling or sink to the floor?
3. What is the volume in liters of a kilogram of ice at 0°C?
4. What is the mass of a bar of aluminum measuring 1.0 cm by 1.0 cm by 10.0 cm?

CHAPTER 2: Substances vs. Mixtures, Chemical vs Physical Change; Law of Conservation of Mass

3. Classify the following as elements, compounds, or mixtures.
   a. table salt **b.** water **c.** iron **d.** stainless steel

SECTION 2.4 CHEMICAL REACTIONS

1. Which one of the following is a chemical change?
   a. Gasoline boils.
   b. Oxygen is added to gasoline.
   c. Gasoline burns.
   d. Gasoline is poured into a tank.
2. Classify each of the following changes as physical or chemical.
   a. A puddle is dried by the sun.
   b. A dark cloth is faded by sunlight.
   c. Bread is toasted.
   d. Soap is mixed with water.
3. Carbon dioxide plus water yields carbonic acid.
   a. Name the product(s) of this reaction.
   b. Name the reactant(s) of this reaction.
4. If 44 grams of carbon dioxide react completely with 18 grams of water, what is the mass of carbonic acid formed?
5. In an engine, octane combines with oxygen to form carbon dioxide and water. If 22.8 grams of octane combine completely with 80 grams of oxygen to form 70.4 grams of carbon dioxide, what mass of water is formed?
6. What is the name of the chemical law on which problems 4 and 5 are based?