

Chemical Quantities

mole : a representative unit
to measure number
of atoms or molecules.

$$6.022 \times 10^{23} \text{ atom} = 1 \text{ mole}$$

Molar mass - grams/moles

Atomic mass = Molar Mass

Na - 22.99 g/mol

H - 1.0079 g/mol

As - 74.92 g/mol

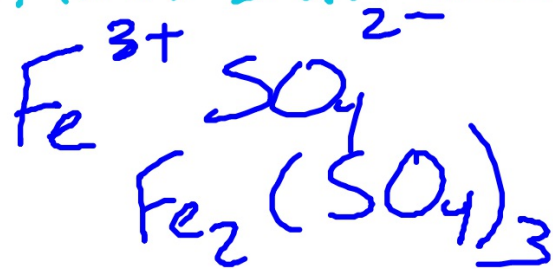
Formula molar mass



$$\text{Na} - 22.99 \times 1 = 22.99$$

$$\text{Cl} - 35.45 \times 1 = \frac{35.45}{58.44} \text{ g/mol}$$

Iron (III) Sulfate



$$\text{Fe} - 55.85 \times 2 = 111.7$$

$$\text{S} - 32.07 \times 3 = 96.21$$

$$\text{O} - 16.00 \times 12 = 192.0$$

$$399.91 \text{ g/mol}$$

Sucrose $C_{12}H_{24}O_{12}$

$$C - 12.01 \times 12 = 144.12$$

$$H - 1.0079 \times 24 = 24.1896$$

$$O - 16.00 \times 12 = \frac{192.00}{360.3 \text{ g/mol}}$$

