Ionic Bonding Practice

Name: Period: Notebook:

Rules for Naming Compounds

- ✓ The name for positive ion is written first.
- ✓ The negative ion is written last. And the ending of the element name is changed to "íde."
- ✓ If either ion is a polyatomic, the polyatomic name is used.

Practice 1 (just elements)

lons to bond		Formula Compound Name	
Na	S	Na ₂ S	Sodium Sulfide
К	CI		
AI	I		
Ca	CI		
AI	0		
К	I		
Са	0		
Mg	S		
AI	Ν		
К	N		
Ca	S		
Na	N		
К	0		
Mg	N		
К	S		
Mg	CI		
AI	S		

Practice 2 (some polyatomic ions)

What is a polyatomic ion? Give a few examples:

Use your lonic Bonding Buddy to determine the ions involved in the following bonds. Use the criss-cross method (& shapes if needed) to determine the compound's chemical formula.

	lons		Formula
Magnesium hydroxide	Mg ⁺²	OH ⁻	Mg(OH)₂
potassium hydroxide			
magnesium bromide			
rubidium oxide			
aluminum sulfate			
aluminum phosphate			
sodium hydroxide			
calcium fluoride			
hydrogen sulfide			
sodium chromate			
sodium phosphate			
calcium hydroxide			
calcium iodide			
lithium dichromate			
calcium sulfate			
beryllium nitrate			
strontium bicarbonate			
* copper(II) chloride			
* iron(II) oxide			
* iron(III) oxide			

Practice 3 (naming compounds containing polyatomic ions)

Ba(OH) ₂	Na(NO ₃)
KClO ₃	NH₄F
NH₄OH	Mg(HCO ₃) ₂
Ba(CO ₃)	Ca(HSO ₄) ₂
K(C ₂ H ₃ O ₂)	* Cu(ClO ₃)